## **IPRO 346**

### Objective:

Choose and design a desulfurization method to conform with EPA regulations and economically remove sulfur from coal.

EPA Standard: 81% total sulfur removed
Our Design: 86% of total sulfur removed

**Costing Basis:** 

554,400 tons of coal processed / year

Buying Price:

\$29.00 / ton of coal

\$18.68 / ton of lime

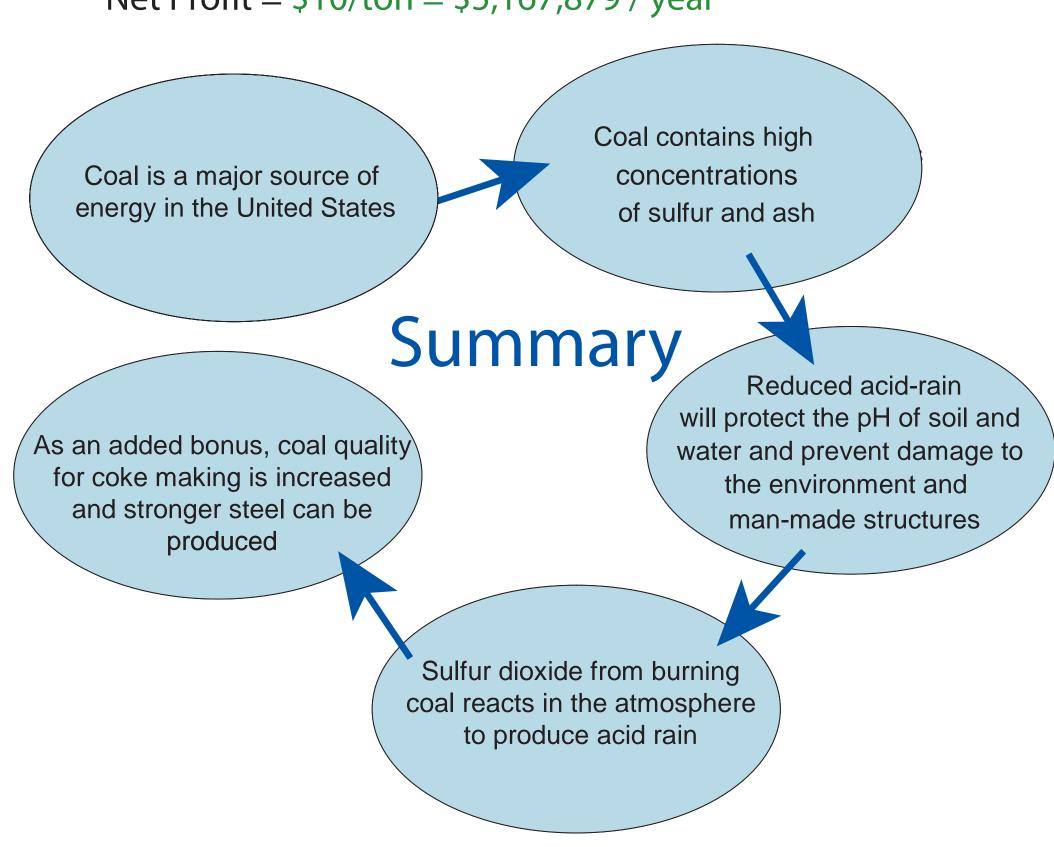
Sale Prices: \$90.58 / ton of cleaned coal

\$23.00 / ton of gypsum

Profit:

Expenses: \$88/ton
Total Sales: \$98/ton

Net Profit = \$10/ton = \$5,167,879 / year



### **Coal Composition**

Sample	Ash	Hydrogen	Carbon	Nitrogen	Oxygen	Sulfur
Number	%	%	%	%	%	%
Illinois #6 <i>Average</i>	11.68	5.04	65.475	1.34	12.655	

### **Sulfur Composition**

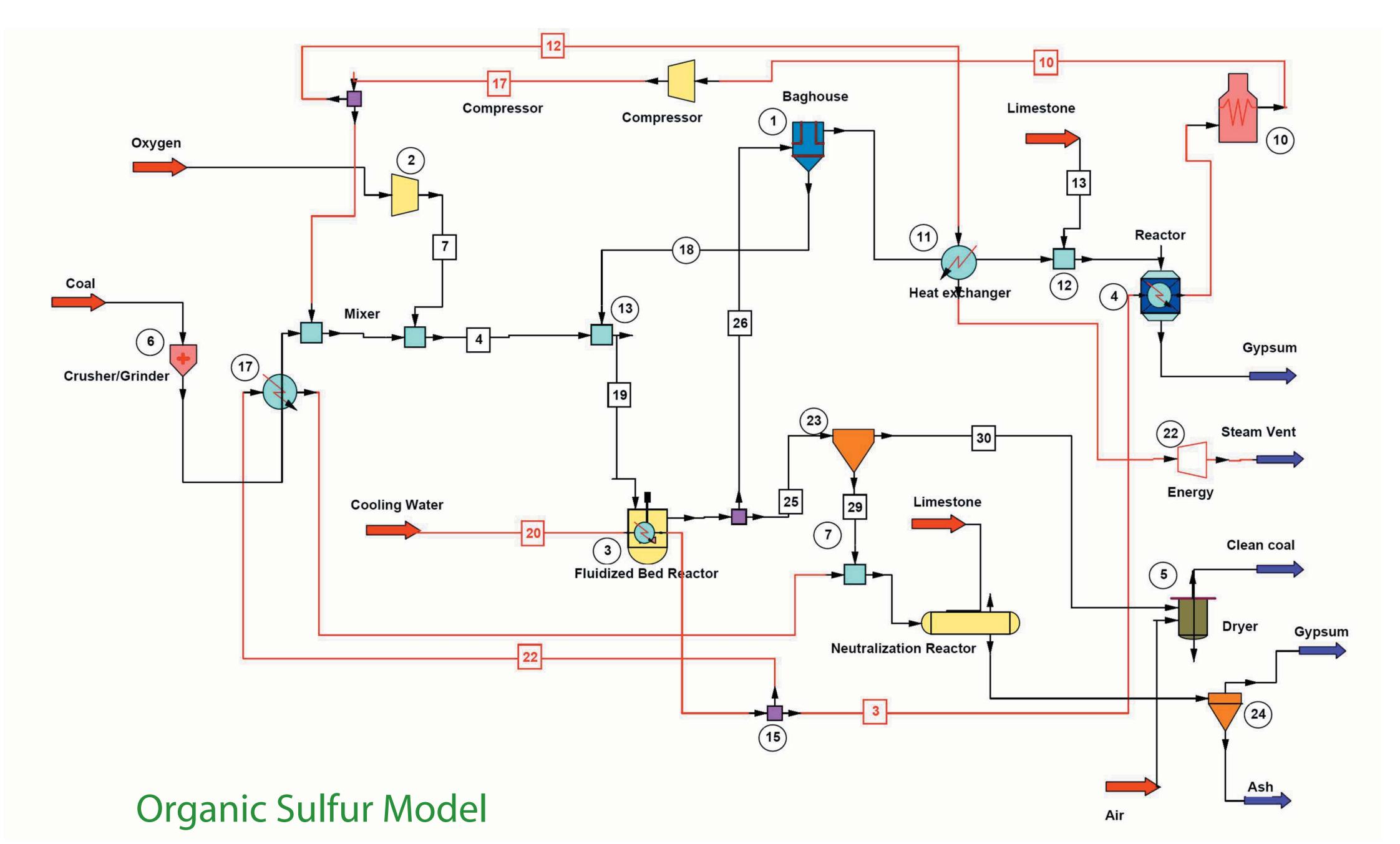
Sample		Elemental Sulfur %	Pyritic %	Organic %	
	Illinois #6 <i>Average</i>	0.175	1.825	1.805	

### **Fluidized Bed Calculation Results**

	Variable	Initial value	Final value
	T (s)	0	1800
Elemental Sulfur	Ca kmol/m³	0.034	0.001818
Pyritic Sulfur	Cb kmol/m <sup>3</sup>	0.0973	0.000436
H+	Cc kmol/m <sup>3</sup>	0	0.233877
SO4	Cd kmol/m <sup>3</sup>	0	0.233877

## Coal Desulfurization:

Design of Coal Desulfurization Processes to Improve the Environment



# S

Thiophenic compounds such as this represent over 60% of the Organic Sulfur in Coal

71.76% Organic Sulfur Removed.

### Pyritic Sulfur Reactions

$$FeS_{2} + 2O_{2} \longrightarrow FeSO_{4} + S$$

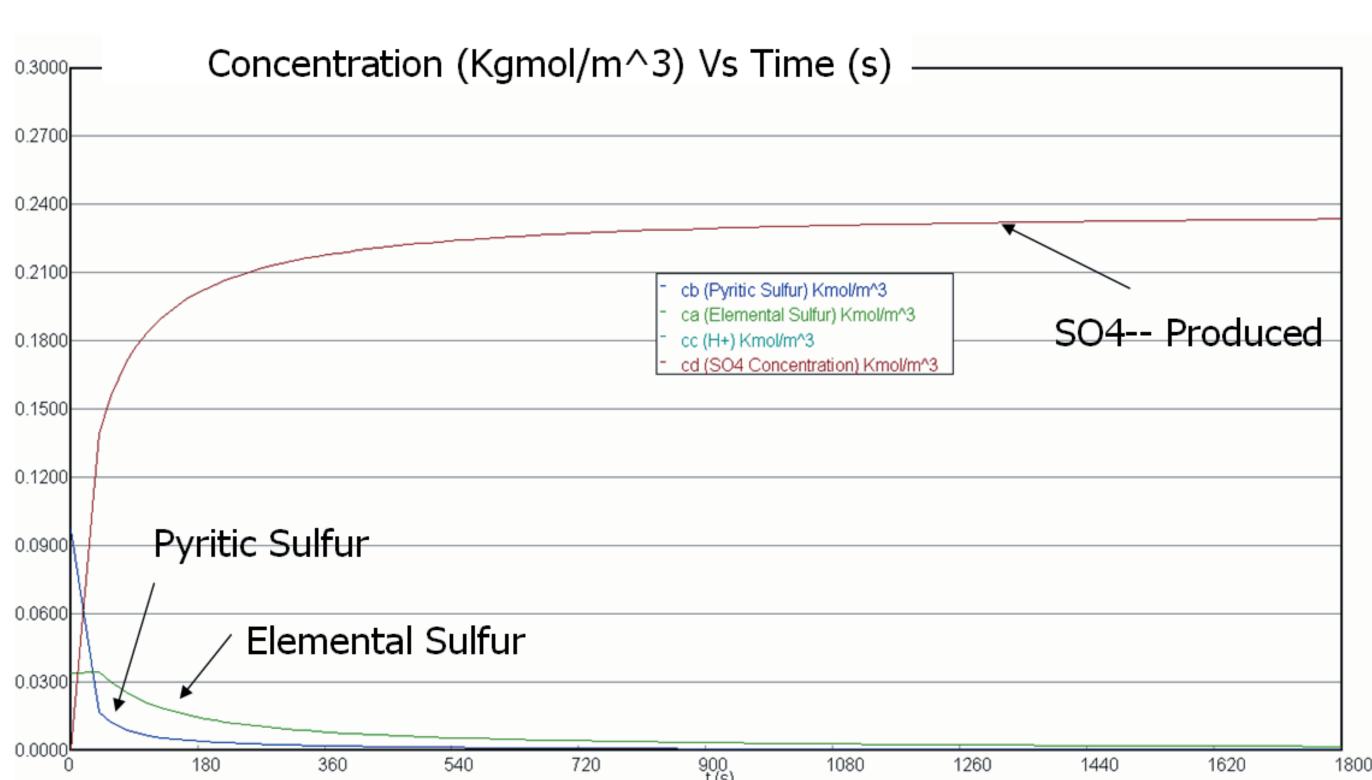
$$FeS_{2} + \frac{7}{2}O_{2} + H_{2}O \longrightarrow Fe^{2+} + 2SO_{4} + 2H^{+}$$

$$FeS_2 + \frac{15}{4}O_2 + \frac{1}{2}H_2O \longrightarrow Fe^{3+} + 2SO_4^{2-} + H^+$$

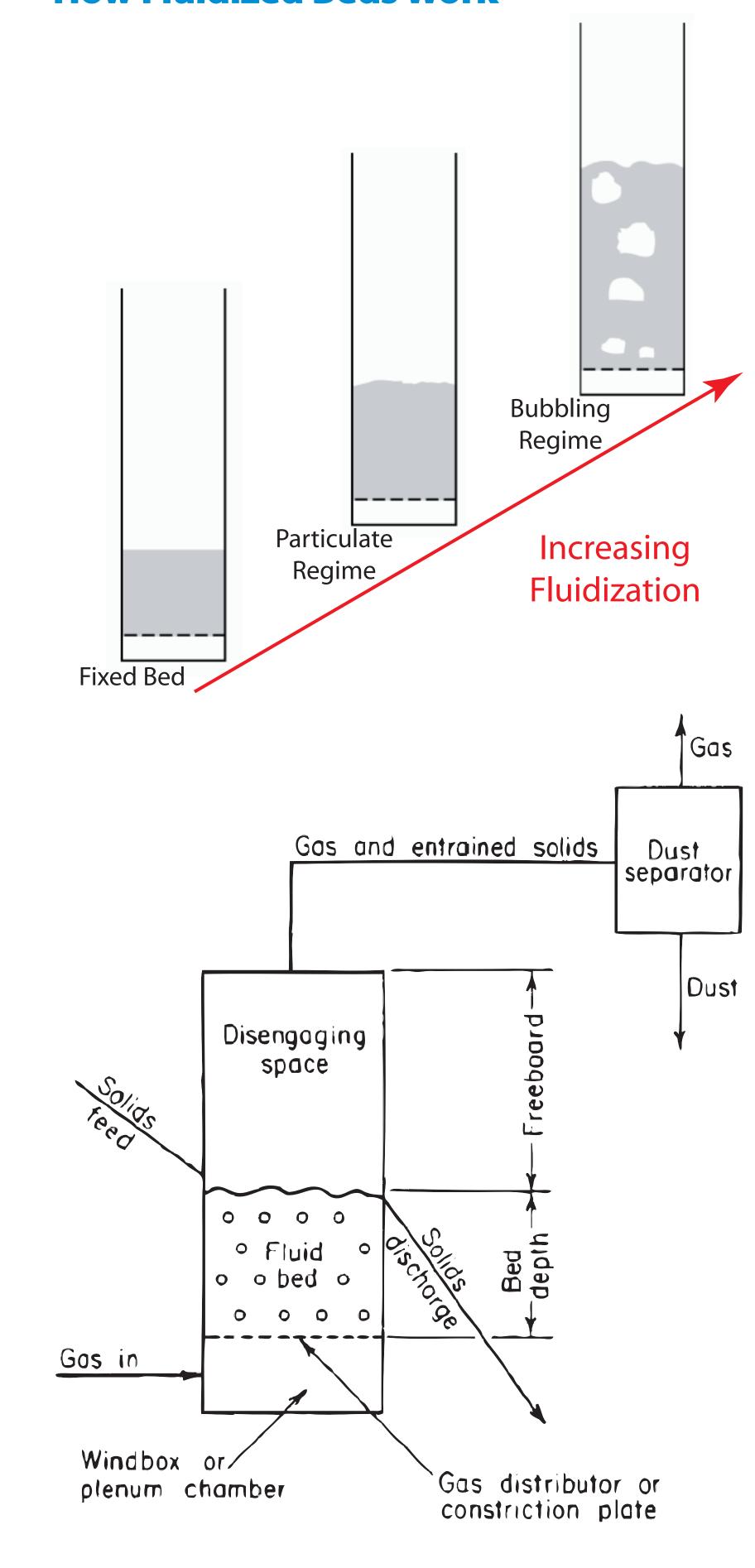
$$FeS_2 + \frac{15}{4}O_2 + 2H_2O \longrightarrow \frac{1}{2}Fe_2O_3 + 4H^+ + 2SO_4^{2-}$$

$$S + \frac{3}{2}O_2 + H_2O \longrightarrow 2H^+ + 2SO_4^{2-}$$

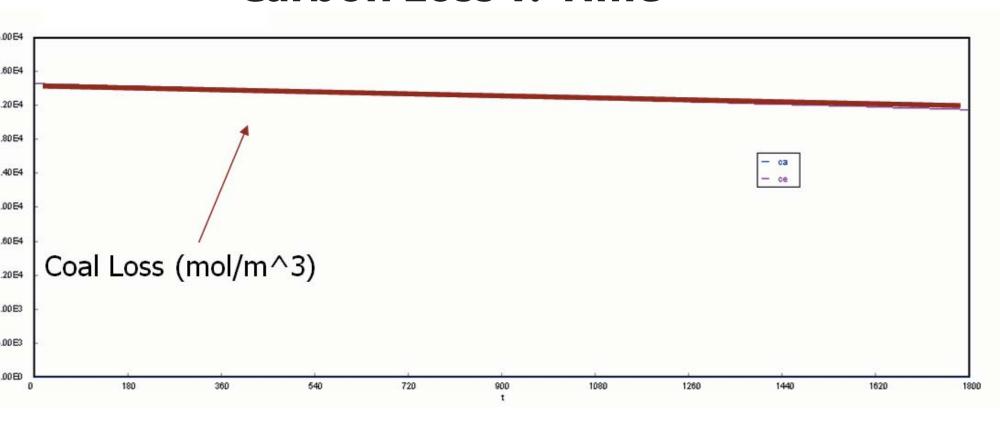




### **How Fluidized Beds Work**



### Carbon Loss v. Time



#### Members:

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